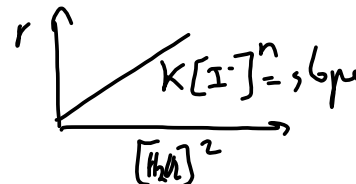
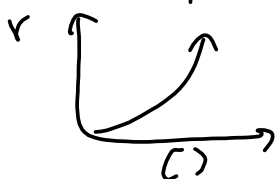
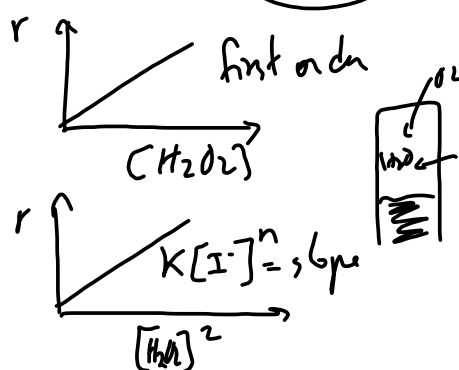
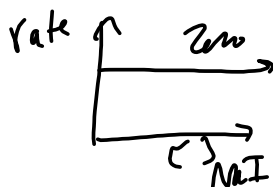


Lab: rate Law.



rate Law:  $r = k [H_2O_2]^m [I^-]^n$

0.12 0.12



$C_1 V_1 = C_2 V_2$

$6\% (8) = C_2 (10)$

$\frac{6\% \times 8}{10} = C_2$

$4.8\% = C_2$

dilution

$4.8\% m/v$

$\frac{4.8 g}{100 mL} H_2O_2$

$C_2$	rate (mL/s)
[ ]	
[ ]	
[ ]	
[ ]	

change to mL/L

$O_2 (g) mL$

density  $\frac{g}{mL}$

$PV = nRT$

$8.314 \text{ kPa} \cdot L / K \cdot mol$

Submit: 2 graphs:

rate versus  $[NaI]$   
rate versus  $[H_2O_2]$   
Graph analysis.

1) table of data collected.

2) table concentration and rate  
one sample calculation